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CHAPTER 9 REVIEW  
Stoichiometry SECTION  
2 PROBLEMS Write the  
answer on the line to  
the left. Show all your  
work in the space  
provided. 1. 4.5 mol  
The following equation  
represents a laboratory  
preparation for oxygen

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### Stoichiometry

gas:  $2\text{KClO}_3 (\text{s}) \rightarrow 2\text{KCl} (\text{s}) + 3\text{O}_2 (\text{g})$  How many moles of  $\text{O}_2$  form if 3.0 mol of  $\text{KClO}_3$  are totally consumed?

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Chapter 9 focuses on reaction stoichiometry: using a balanced chemical equation to calculate the number of grams, moles, or particles of reactants/products

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involved in a chemical reaction. Students...

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Q. 4.3 grams of sodium reacts with 2.6 grams of oxygen to produce 5.9 grams of sodium oxide. What's the limiting reactant?

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### Stoichiometry

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ANSWER Answer the following questions in the space provided.

1. Given the following equation:  $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$

a. What is the value of the coefficient  $x$  in this equation? 4

b. What is the molar mass of  $C_3H_4$ ? 40.07 g/mol

c. What is the mole ratio of  $O_2$  to  $H_2O$ ? 2:1



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Stoichiometry 9-1

Introduction to

Stoichiometry

Composition

Stoichiometry - deals  
with mass relationships

of elements in

compounds Reaction

Stoichiometry -

Involves mass

relationships between

reactants and products

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### Stoichiometry

in a chemical reaction

#### I. Reaction

Stoichiometry

Problems A. Four

problem Types, One

Common Solution

### **Chapter 9 - Stoichiometry**

5. Given the following unbalanced equation:



$\rightarrow \text{NO}_2(\text{g})$  a. Balance

the equation. b. What

is the mole ratio of

$\text{NO}_2$  to  $\text{O}_2$ ? c. If 20.0

mol of  $\text{NO}_2$  form, how

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many moles of  $O_2$   
must have been

consumed? d. Twice as  
many moles of  $NO_2$   
form as moles of  $N_2O$   
are consumed. True or  
False? e. Twice as  
many grams of  $NO_2$   
form as grams of  $N_2O$   
are consumed. True or  
False?

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**Stoichiometry help?**  
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EDIT--Changes must be

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made through “File info” ... Reaction stoichiometry, the subject of this chapter, is based on chemical equations and the law of conservation of mass. All reaction stoichiometry ... The number of significant figures in the answer

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**be made ...**

Chapter 9 Review  
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Answers  
Section 1 Eventually,

you will definitely discover a new experience and ability by spending more cash. yet when? reach you take on that you require to get those all needs when having significantly cash? Why don't you try to get something basic in the beginning?

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### Answers Section 1

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ANSWER Answer the following questions in the space provided. 1. Given the following equation:  $C_3H_4(g) + x \cdot O_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$  a. What is the value of the coefficient . x. in this equation? b. What is the molar mass of  $C_3H_4$ ? c. How many moles are in an 8.0 g sample of  $C_3H_4$ ? 2. a.

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What is meant by .  
ideal conditions

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Stoichiometry b.

Theoretically, how many moles of  $\text{NH}_3$  will be produced?

PROBLEMS Write the answer on the line to the left, Show all your work in the space provided. 1 88% The actual yield of a reaction is 22 g and the theoretical yield is

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25 g. Calculate the percentage yield. 2.

6.0 mol of  $N_2$  are mixed with 12.0 mol of  $H_2$  according to the ...

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SECTION 9.2.

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### Stoichiometry Answers

PROBLEMS Write the answer on the line to the left. Show all your work in the space provided.

1. The following equation represents a laboratory preparation for oxygen gas:  $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$  How many grams of  $\text{O}_2$  form if 3.0 mol of  $\text{KClO}_3$  are totally consumed?

2. Given the following equation

...

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### Stoichiometry

#### **CHAPTER 9 REVIEW**

5.0 g Cu 1 mol Cu 1  
mol Ag 107.9 g Ag =  
8.5 g Ag. 63.5 g Cu 1  
mol Cu 1 mol Cu.  $8.5 \times$   
 $100 = 55.9 \%$  yield.

15.2 CHAPTER 11:  
STOICHIOMETRY. MOLE  
TO MOLE RATIO. When  
nitrogen and hydrogen  
gas are heated under  
the correct conditions,  
ammonia gas (NH<sub>3</sub>) is  
formed. a. RXN: 1. N<sub>2</sub>  
+ 3. H<sub>2</sub> ( 2. NH<sub>3</sub>. b.

**CHAPTER 11:**  
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### **Stoichiometry**

### **STOICHIOMETRY**

At higher pressures, the effect of the volume of the gas molecules themselves on  $Z$  would increase compressibility (see Figure 9.35). (d) Once again, at low pressures, the effect of intermolecular attractions on  $Z$  would be more important than the correction factor for the volume of the gas molecules themselves, though

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perhaps still small.

## Answers

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including balanced  
equations for 3-6.  
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## **Stoichiometry Chapter 9 - Mrs. Baker's Chemistry Website**

9 c pancake mix  $\times$  1  
egg 2 c pancake mix =  
4.5 eggs. Note how the  
units cups pancake mix  
canceled, leaving us

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### Stoichiometry

Answers  
with units of eggs. This is the formal,

mathematical way of getting our amounts to mix with 9 c of pancake mix. We can use a similar conversion factor for the amount of milk: 9 c pancake mix  $\times$  1/2 c milk / 2 c pancake mix = 2.25 c milk

## **Chapter 6 - Stoichiometry and the Mole - CHE**

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Overview of

Stoichiometry Chapter  
Exam Instructions.

Choose your answers to the questions and click 'Next' to see the next set of questions. You can skip questions if you would like and come back ...

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